



# Optical Properties of Natural and Synthetic Minerals

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**Abstract.** The results of investigation of optical absorption and photoluminescence (PL) of topaz, beryl and yttrium aluminium garnet crystals doped with different concentrations of transition ions exposed to fast neutron irradiation and electron irradiation are presented. We suppose that irradiation leads to the formation of two types of complex centers: “Me<sup>2+</sup>- F<sup>+</sup> (or F) centre” and complex centers, which consist of a cation vacancy and an impurity (iron, manganese and chromium) ion. Exchange interaction between radiation defects and impurity ions during neutron or electron irradiation gives rise to appearance of additional absorption and luminescence band broadening in investigated crystals.

**Keywords** – inorganic compounds, absorption spectra, photoluminescence, neutron irradiation.

## I INTRODUCTION

Natural and synthetic minerals such as topaz, beryl and garnet are used not only as gemstone but have a practical application. Topaz crystal is a suitable material for dosimetric applications [1]-[3]. Beryl and garnet crystals doped with transition metal ions find application as laser materials [4], [5]. Most gemstone crystals have an intrinsic color due to the optical absorption and presence of color centers that are produced by impurities like transition metals and elements of Earth-rare, or defects in the crystalline lattice [6]. Defect and disorder creation has been studied in detail, but the renewed interest arises from the use of electron and ion beams for synthesis and modification of materials.

Potential application of this material to a great extent is determined by its optical properties, which in turn are determined not only by structure but also by presence of defects. Therefore it is important to estimate the present material defects and their effect on properties of material.

The paper presents results of the effects of neutron and electron irradiation on optical properties of natural (topaz, beryl) and synthetic (beryl, yttrium aluminium garnet) minerals.

## II MATERIALS AND METHODS

Oxide materials Y<sub>3</sub>Al<sub>5</sub>O<sub>12</sub>, natural and synthetic beryl Be<sub>3</sub>Al<sub>2</sub>Si<sub>6</sub>O<sub>18</sub>, topaz Al<sub>2</sub>[SiO<sub>4</sub>](F, OH)<sub>2</sub> have been used for research.

The investigated YAG samples were cut from single crystals grown by the Czochralski technique, crystal had the composition Y<sub>3</sub>Al<sub>5-x</sub>Mn<sub>x</sub>O<sub>12</sub>, where x = 0.012, 0.025, 0.037 and 0.05. Samples of beryl were obtained from variety of sources: natural crystals from Ural Mountains and synthetic grown by flux method in Russia. Natural colourless and blue topaz from

Ukraine was investigated. The sample thickness was in the range of 0.1–2.15 mm.

The neutron irradiation was performed at the Latvian 5 MW water-water research reactor. The fluence of fast neutrons with energy > 0.1 MeV was in the range 10<sup>14</sup>–5·10<sup>18</sup> cm<sup>-2</sup>. A cadmium filter was used for absorption of thermal neutrons. Irradiation by electrons with energy of 3.5 MeV of integral fluence up to 10<sup>16</sup> cm<sup>-2</sup> was carried out in accelerator ELU (Salaspils, Latvia). The sample temperature did not exceed 60°C during the electron irradiation and was below 40°C during the neutron irradiation.

The luminescence studies were carried out using two setups for spectral measurements. The first setup was equipped with a SPM-2 monochromator having diffraction grating of 651 lines per mm. The crystals were excited with a high pressure xenon lamp (DKSEL-1000) 1000 W connected to the monochromator SPM-1 (Carl Zeiss Iena) with a quartz prism. Signal detection was carried out with a photomultiplier tube (FEU-119) by a synchronic detection method. The second setup was equipped with a deuterium lamp LDD-400 as a source of UV light and a grating monochromator MDR-2 in the excitation channel. In the case of PL measurements the luminescence signal was analyzed by Shamrock monochromator SR-303i-B using a grating with 150 lines/mm and 500 nm blaze and detected with a CCD camera DV420A-BU2 in the accumulation mode. The standard technique to measure absorption spectra has been based on a “Specord M-40” two-beam spectrometer operating in the 200-900 nm regions and “Specord 210” (Analytikjena) double-beam spectrophotometer operating in the spectral region of 190-1100 nm. Optical measurements before and after irradiation were carried out in the temperature range 8-300 K.

### III RESULTS AND DISCUSSION

#### A. Topaz crystals

Topaz is an aluminium fluorosilicate with a fairly constant chemical composition  $\text{Al}_2\text{SiO}_4(\text{OH},\text{F})_2$ . The only major variation found in different samples is related to the  $[\text{OH}]/[\text{F}]$  concentration ratio. The structure of topaz consists of  $\text{SiO}_4$  groups linking octahedral chains of  $\text{Al}[\text{O}_4(\text{F},\text{OH})_2]$  in a zigzag fashion parallel to the crystalline c-axis. Four of six anions surrounding  $\text{Al}^{3+}$  ion belong to  $\text{SiO}_4$  tetrahedron and the remaining two anions are  $\text{F}^-$  or  $\text{OH}^-$  group. Topaz crystallizes in the orthorhombic system, space group Pbnm [7], [8] and is normally found as well-developed prismatic crystals with pyramidal terminations

Natural topaz stone has the structure of an aluminum silicate fluoride hydroxide  $\text{Al}_2[\text{SiO}_4](\text{F},\text{OH})_2$ , and is usually colorless or has an unattractive color. Figure 1. gives the absorption spectra of topaz single crystals before irradiation (curve 1), after fast neutron irradiation (curve 2) and natural blue topaz from Volodarsk Volynskii, Ukraine (curve 3). Before irradiation some zero-phonon lines were observed in the region of the  $\text{Cr}^{3+}$  ion electron transition  ${}^4\text{A}_{2g} \rightarrow {}^2\text{E}_g$  in topaz spectra (Fig. 2). After irradiation in absorption spectra there appears an intense absorption at 230 nm and occur the characteristic maxima at 305, 410, 446, ~590 nm. Two lines around 230 and 305 nm belong to a single defect such as silanone ( $=\text{Si}=\text{O}$ ) as it was suggested in [9] basing on similar lines observed in spodumene. Blue color is connected with a broad absorption band in the red part of the spectrum generated by the so-called R-centers (two F-vacancies with two trapped electrons) [10]. It was previously concluded that the optical absorption band centered at 620 nm is closely correlated with an  $\text{O}^-$  center interacting with two Al ions of the topaz structure. This  $\text{O}^-$  center is produced by the irradiation in the hydroxyl sites which substitutes for fluorine in the topaz structure [11], [12]. As suggested Priest V. at al [13] the blue defects in neutron damaged topaz is the doubly occupied dangling silicon bond. We assume that the band ~ 620 nm includes not only the above mentioned bands, but also the bands associated with the presence of impurities of  $\text{Cr}^{3+}$ ,  $\text{Fe}^{2+}$  and  $\text{Mn}^{2+}$  ions. The calculations performed in [14] using  $\text{AlCrFH}_7\text{O}_{11}\text{Si}$  model molecule with unrestricted CIS method to simulate UV and visible spectra shows that the theoretical spectrum of the model is in a considerably good agreement with experimental one.

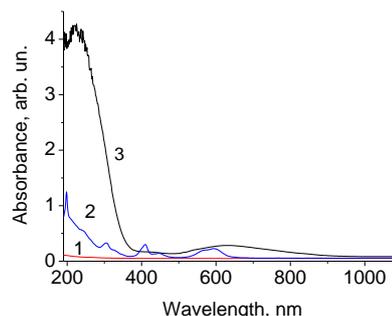
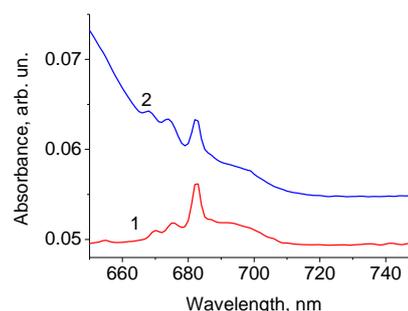


Fig.1. Topaz absorption spectra before (1) and after neutron irradiation (2), (3) natural blue topaz from Volodarsk Volynskii,



Ukraine.

Fig. 2. Topaz absorption spectra before (1) and after neutron irradiation (2)

Photoluminescence spectra in natural colourless topaz excited at  $\lambda_{\text{exc}} = 251$  nm and  $\lambda_{\text{exc}} = 447$  nm registered at 300 K are shown in Fig. 3. In photoluminescence spectra of topaz crystal excited at 251 nm, it is possible to separate an intensive band 350 nm with shoulder about 364 nm, and less expressed 390 nm, 420 nm and 464 nm bands. Upon excitation 447 nm it is observed two wide bands at maxima at 575 and 800 nm.

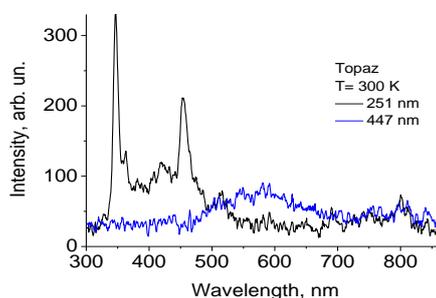


Fig.3. Photoluminescence spectra of topaz crystal at 230 and 447 nm excitation wavelengths.

Luminescence at 390 nm belongs to own  $[\text{SiO}_4]^{3+}$  hole centers. The emission band with maxima 464 nm is connected with  $[\text{AlO}_4]^{4-}$  center arising from isomorphic substitution of silicon by aluminum in the silicon-oxygen tetrahedron. Wide luminescence bands

at excitation wavelength 447 nm may be connected with presence of  $Mn^{2+}$  (~ 600 nm) and  $Fe^{3+}$  (~ 800 nm) isomorphically substituting for  $Si^{4+}$  in the silicon-oxygen tetrahedron.

### B. Beryl crystal

Beryl ( $Be_3Al_2Si_6O_{18}$ ) is a silicate mineral belonging to the cyclosilicate class. It has a honeycomb-like hexagonal structure and belongs to the P6/mcc space group. The axial parameters are  $c=9.17 \text{ \AA}$  and  $a=9.21 \text{ \AA}$  [15, 16]. Each ion of Al is surrounded by 6 oxygen atoms located in the vertices of almost regular octahedron. Each octahedral complex  $AlO_6$  is linked to six  $SiO_4$  tetrahedron through common oxygen atoms. The disposition of the octahedral is such that allows the formation of  $Si_6O_{18}$  rings. Each beryllium is surrounded by 4 oxygens forming distorted tetrahedron. The  $Si_6O_{18}$  rings lay one above the other along the C6 axis forming intercommunicating channels of quite large dimensions. The oxygen that is shared among  $Si^{4+}$  ions is an O1 type oxygen and the oxygen that is linked to  $Al^{3+}$  and  $Be^{2+}$  ions is an O2 type oxygen. Taking into account this configuration, each  $Si^{4+}$  ion is surrounded by two O1 oxygens and two O2 oxygens. The diameter of the channels varies from a minimum of  $2.8 \text{ \AA}$  in the plane of the  $Si_6O_{18}$  ring and a maximum of  $5.1 \text{ \AA}$ , midway between two neighboring rings. Because of this structure, impurities can be substitutional, interstitial and also located in the channels.

The color of beryl  $Be_3Al_2Si_6O_{18}$  is usually determined by its  $Fe^{2+}$  and  $Fe^{3+}$  content.  $Fe^{2+}$  produces pale blue color,  $Fe^{3+}$  produces golden-yellow color, and when both  $Fe^{2+}$  and  $Fe^{3+}$  are present, the color is dark blue. Green color in iron-containing beryl usually results from a mixture of blue and yellow components. Green color can also come from  $Cr^{3+}$  as it is found in various emeralds. Beryl is naturally transparent, however inclusions and impurities may make it opaque.

Absorption spectra for three beryl crystals are shown in Fig. 4. Beryl, when containing sufficient chromium (about 0.1% or more) is characterized by two broad bands near 430 and 600 nm (Fig. 4, curve 1), as well sharp R-lines (Fig 5.)

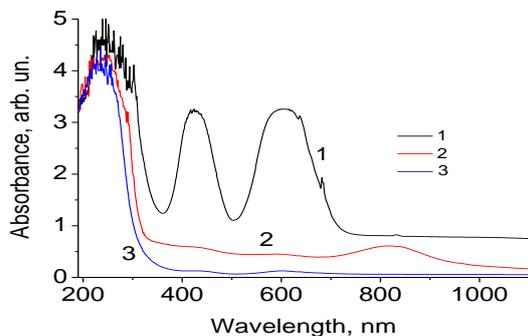


Fig.4. Absorption spectra of beryl crystals: 1. synthetic beryl containing Cr 0.1% chromium impurity, 2. natural beryl from Ural, 3. synthetic flux beryl.

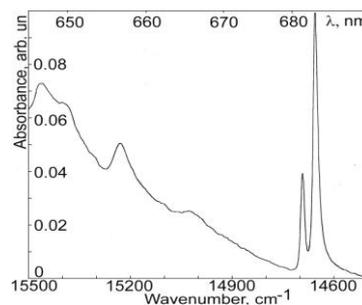


Fig. 5. Absorption spectra of beryl containing 0.1 %Cr:

The observed absorption bands associated with electronic transitions of chromium ion: 630 nm transition  ${}^4A_{2g}(F) \rightarrow {}^4T_{2g}(F)$  and 431 nm transition  ${}^4A_{2g}(F) \rightarrow {}^4T_{1g}(F)$  [17]. The wide absorption band with peak in the near-infrared range at 813 nm is observed in natural beryl spectra. The band is generally ascribed to internal electron transition of  ${}^5T_2({}^5D) \rightarrow {}^5E({}^5D)$  of  $Fe_{VI}^{2+}$  ions localized in octahedral aluminum sites of beryl [18, 19]. Absorption spectra of beryl containing chromium before and after neutron irradiation are given in Fig.6.

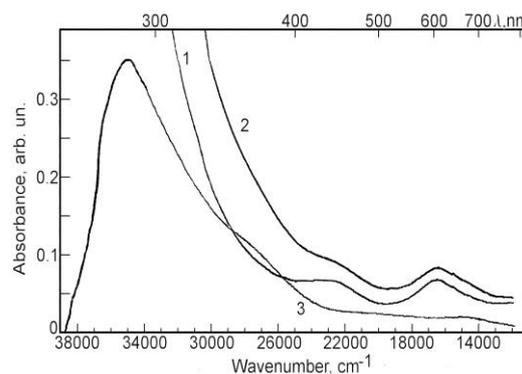


Fig. 6. Absorption spectra of beryl:Cr: 1. before irradiation, 2. after fast neutron irradiation  $\Phi=10^{16} \text{ cm}^{-2}$ , 3. additional absorption spectrum.

Fast neutron irradiation produces additional bands with maxima: 286, 370, 500 and 667 nm. Most probably these bands are due to anion intrinsic defects (the  $F^+$ ,  $F$ ,  $F_2$ ,  $F_2^+$  and  $F_2^-$  centers) similar to observed in  $Al_2O_3$  [20], [21]. Natural pale blue beryl crystals contain iron ions and small amount of chromium ions. Prior to irradiation there were not observed absorption bands associated with electronic transitions  ${}^4A_{2g}(F) \rightarrow {}^4T_{2g}(F)$  (630 nm) and  ${}^4A_{2g}(F) \rightarrow {}^4T_{1g}(F)$  (431 nm) of chromium ions. Absorption spectra contain only band 813 nm associated with  $Fe^{2+}$  ions. After the impact of fast neutron fluence of  $10^{13}$ - $10^{17} \text{ cm}^{-2}$  the intensity of 813 nm band reduced and the absorption edge shifted due to appearance of an intense absorption band in the UV region. Besides, after irradiation there appears the additional band with maximum 690 nm (Fig. 7). The crystal color is changed from pale blue to light green.

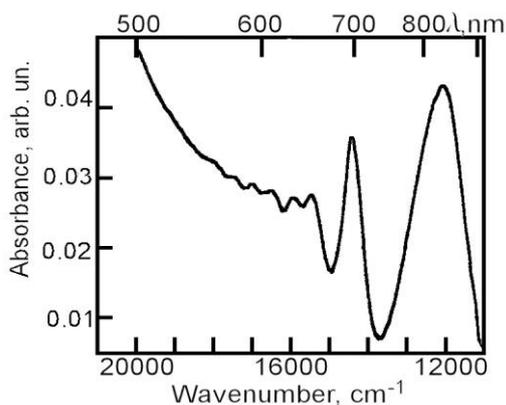


Fig.7. Absorption spectra of natural pale blue beryl after fast neutron irradiation  $\Phi = 10^{16} \text{cm}^{-2}$ .

Dependence of intensity of 813 nm (1) and 694 nm (2) absorption bands on fast neutron fluence in pale blue beryl is shown in Figure 8.

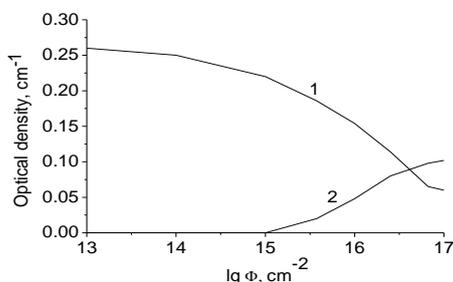
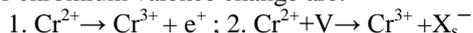


Fig. 8. Pale blue beryl absorption band 813 nm (1) and 694 nm (2) intensities as functions of fast neutron fluence.

As previously reported [22]-[24] irradiation produces radicals  $\text{NO}_3^0$  and  $\text{CO}_3^-$ , which are called Maxixe and Maxixe-type color centers and which seem to be related to blue colors in beryl. We suppose that the band with maxima 690 nm belongs to complex center which consists from  $\text{Cr}^{3+}$  ions and radiation defects (F or  $\text{F}^+$ - center). Concentration of F and  $\text{F}^+$  centers are increasing with neutron fluence increasing. The number of perturbed chromium sites also increased and the band became more intense. The containing chromium impurity ions may be the chromium ions in the lower valence, such as bivalent  $\text{Cr}^{2+}$  impurity ions. The  $\text{Cr}^{2+}$  impurity ions are oxidized into  $\text{Cr}^{3+}$ . A  $\text{Cr}^{3+}$  impurity ion has the same valence as the substituted  $\text{Al}^{3+}$  ion. Possible processes for chromium valence change are:



( $\text{X}_s^-$  -the constituent anions or radical in normal lattice of beryl crystal [25]). Therefore, the newly produced  $\text{Cr}^{3+}$  impurity ions prefer to lie in the octahedral sites in order to maintain the local charge neutrality of the crystal.

The photoluminescence spectra of flux beryl at  $\lambda_{\text{ex}} = 271 \text{ nm}$  and different temperature are given in Fig. 9. Wide band at 740 nm observed at  $T = 300 \text{ K}$  is connected with  $\text{Fe}^{2+}$  ions. Narrow lines in region from 680 to 720 nm at  $T = 8 \text{ K}$  belong to single  $\text{Cr}^{3+}$  (R-

lines) and  $\text{Cr}^{3+}$ - pairs (N- lines). The photoluminescence band at 525 nm intensity increases with temperature decrease.

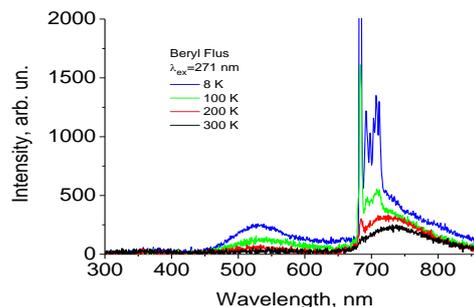


Fig.9. Photoluminescence spectra of beryl crystal at different temperature, excitation wavelengths 271 nm.

### C. Yttrium aluminum garnet

The cubic cell of garnet compound  $\text{A}_3\text{B}'_2\text{B}''_3\text{O}_{12}$  contains eight formula units, where A, B', B'' are metal ions occupying sites of different symmetry [26]. The crystal structure of  $\text{Y}_3\text{Al}_5\text{O}_{12}$  (YAG) has a bcc structure (space group  $\text{Ia}\bar{3}\text{d}$  or  $\text{O}_h^{10}$ ) with 160 atoms in the cubic (primitive) cell. Y ions (A atoms) occupy 24 (c) sites and each of them is dodecahedrally coordinated to eight oxygen ions. Oxygen atoms occupy 96 (h) sites whose exact location depends on three structural parameters - x, y, and z and are different for different garnet oxides. There are two different sites for Al:  $\text{Al}_{\text{oct}}$  (B' atoms) occupy 16 (a) sites with octahedral point symmetry ( $\text{C}_{3i}$ ); and  $\text{Al}_{\text{tet}}$  (B'') atoms occupy 24 (d) sites with tetrahedral point symmetry ( $\text{S}_4$ ). The garnet structure may be considered as interconnected octahedrons, tetrahedrons, and dodecahedrons with shared oxygen atoms at the corners [26]. Each oxygen ion is a member of two dodecahedrons, one octahedron, and one tetrahedron.  $\text{Y}^{3+}$  and  $\text{Al}^{3+}$  in YAG can be replaced by many kinds of other cations with different valence and size within a suitable range.

The optical absorption spectra of four garnet single crystals measured at room temperature before irradiation are similar. Figure 10. gives the absorption spectra of yttrium aluminum garnet single crystals (curve 1) before irradiation, after fast neutron irradiation (curve 2) and after electron irradiation (curve 3). Before irradiation there are observed bands at 370, 500 and 830 nm. After neutron irradiation additional bands with maxima 549.5, 595.2 and 714.3 nm appear in absorption spectra of YAG containing manganese impurities. The 370 and 500 nm bands intensity increases but the 830 nm band intensity decreases after fast neutron irradiation. A weak band 370 nm corresponds to F-type centers [27].

Broad orange luminescence band in the range 580-620 nm appears at room temperature (Fig. 11, curve 1). Luminescence spectrum of  $\text{Y}_3\text{Al}_{5-x}\text{Mn}_x\text{O}_{12}$  single crystal measured at liquid nitrogen temperature contains the following main features: a sharp line at

587 nm and a series of weaker lines on the long wavelength side of the sharp lines (curve 2, Fig.11).

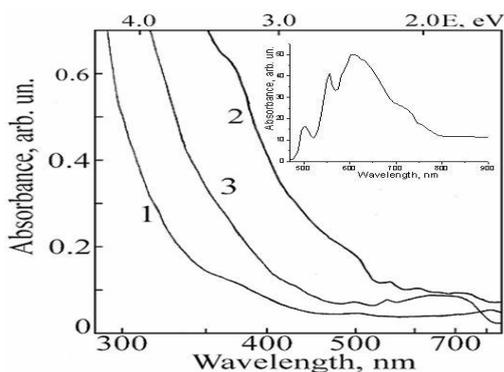


Fig. 10. . Absorption spectra of  $Y_3Al_5Mn_{0.12}$ : 1. before irradiation, 2. after neutron irradiation  $\Phi = 10^{16} \text{cm}^{-2}$ , 3. after electron irradiation  $\Phi = 10^{16} \text{cm}^{-2}$ . In the inset: additional absorption spectrum of neutron irradiated  $Y_3Al_{4.963}Mn_{0.037}O_{12}$  crystal.

After irradiation with fast neutrons at fluence  $6 \cdot 10^{18} \text{cm}^{-2}$  the photoluminescence spectrum at liquid nitrogen temperature consists of a broad structureless band with a maximum at 607 nm (Fig. 11, inset). A little sharp line is seen also as a shoulder of the main band. Neutron irradiation causes production of anion and cation vacancies in the crystal. F-centre absorption band at 370 nm arises. After PL and optical measurement thermal annealing of the  $Y_3Al_{4.988}Mn_{0.012}O_{12}$  and  $Y_3Al_{4.975}Mn_{0.025}O_{12}$  crystals has been carried out. The crystals were heated in air at 473 K during 2 hours. The PL spectra measured after annealing at liquid nitrogen temperature are similar to the PL spectra before irradiation.

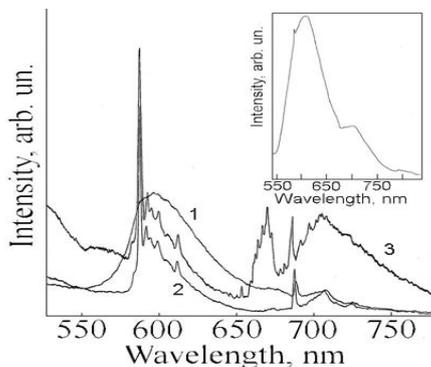


Fig. 11. Photoluminescence spectra of  $Y_3Al_{5-x}Mn_xO_{12}$  at  $\lambda_{ex}=470 \text{ nm}$ : 1. at room temperature, 2. at 80 K, and 3. after electron irradiation. The inset shows photoluminescence spectrum of a neutron- irradiated crystal.

It is suggested, that the reversibility of broadening of luminescence lines of  $Mn^{2+}$  ions is the evidence of the structure ordering reconstruction after annealing. The relatively low annealing temperature shows, that the structure changes after neutron irradiation concern only cation sublattice. Photoluminescence spectra of  $Y_3Al_{4.988}Mn_{0.012}O_{12}$  carried out at liquid nitrogen temperature after electron irradiation have discovered a series of relatively narrow lines in 645-680 nm spectral regions.  $Mn^{2+}$  ions have a  $3d^5$  electron configuration. All electron transitions for this

configuration are forbidden by the selection rules of spin and parity. The restrictions caused by the selection rule are removed owing to the exchange interaction. The emission with maxima around 590 nm is attributed to the  $4T_1g \rightarrow 6A_1g$  transition of  $Mn^{2+}$  ions in octahedral sites. Luminescence spectrum of  $Mn^{4+}$  ( $3d^3$ ), which is isoelectronic to  $Cr^{3+}$  consists of the sharp R lines peaked around 645-675 nm, which are caused mainly by the  $2E_g \rightarrow 4A_2g$  transitions and partly by the  $4T_2g \rightarrow 4A_2g$  transitions. These bands are observed after electron irradiation of  $Y_3Al_{5-x}Mn_xO_{12}$  together with the  $Mn^{2+}$  ion luminescence in octahedral sites. In contrast to paper [5] the green luminescence was not observed in our single crystals. The broad band emission of  $Cr^{3+}$  with a maximum at 695 nm is ascribed to the  $4T_2g \rightarrow 4A_2g$  transition, the sharp and narrow bands at 685 nm are due to  $2E_g \rightarrow 4A_2g$ . The diffused band centered at 675 nm (anti-Stokes vibronic sidebands) and two sharp diffused bands at 710 and 725 nm corresponding to the vibronic sidebands are observed in all investigated YAG crystals before and after irradiation (Fig.10). Hodges et al. [28] pointed out that  $Mn^{2+}$  in yttrium aluminum garnet can enter into any of three kinds of sites which are located in the center of dodecahedron, octahedron and tetrahedron, respectively [29, 30]. The presence of a nearby F or  $F^+$  center can modify the neighborhood the  $Mn^{2+}$  impurity site to produce an effective increase in the crystal field. In addition to the shift in energy, the perturbation induced by the F or  $F^+$  center can increase the oscillator strength by introducing odd components to the crystal field, or by exchange coupling between the F or  $F^+$  center and the  $Mn^{2+}$  impurity. The enhancement of the oscillator strength by the exchange interaction has been previously observed between F centers and transition metal impurity ions in  $MgF_2$ ,  $MgAl_2O_4$  and  $MgO$  [31, 32]. As the concentration of F and  $F^+$  centers increased, the number of perturbed manganese sites also increased, so that the excitation peaks became more intense as a function of neutron dose. It is possible to assume that the band broadening is connected with a complex center " $Mn^{2+}F^+$  (or F centre)". A part of holes generated after irradiation are trapped by  $Mn^{2+}$ , which then produces  $Mn^{3+}$  and  $Mn^{4+}$  ions.

#### IV CONCLUSIONS

The absorption and luminescence spectra behaviour of natural (topaz, beryl) and synthetic (beryl, yttrium aluminum garnet) minerals containing transition metal ions exposed to a fluence of fast neutrons up to  $10^{18} \text{cm}^{-2}$  ( $E > 0.1 \text{ MeV}$ ) and electron are investigated.

We assume that the band  $\sim 620 \text{ nm}$  in topaz crystals is associated with the presence of impurities of  $Cr^{3+}$ ,  $Fe^{2+}$  and  $Mn^{2+}$  ions. Band with maxima 690 nm in beryl crystal belongs to complex center which consists from  $Cr^{3+}$  ions and radiation defects (F or  $F^+$  center). Band broadening in aluminium yttrium garnet is connected with a complex center " $Mn^{2+}F^+$  (or F centre)". Electron irradiation produced broad band with a complex structure related to  $Mn^{4+}$  ions.

Exchange interaction between radiation defect and impurity ions during neutron irradiation and electron irradiation leads to appearance of additional absorption and luminescence bands and bands broadening in investigated crystals.

#### V ACKNOWLEDGMENTS

This work was supported by the grant of the Latvian Government (No 402/2012).

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